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Uses of Common Salt :

(1)

- \* It is used as a raw material for making useful chemicals e.g. NaOH (Caustic Soda)  $\text{Na}_2\text{CO}_3$  (Washing Soda),  $\text{NaHCO}_3$  (Baking Soda),
- \* In cooking food. It improves the flavour of food.
- \* It is required by our body for the working of nervous system.
- \* It is used as a preservative in pickles meat & fish.
- \* It is used in the manufacture of Soap.
- \* It is used to melt ice which collects on the road in cold countries.

Chemicals From Common SaltSodium Hydroxide:

Sodium hydroxide is commonly known as caustic soda. The chemical formula of sodium hydroxide is  $\text{NaOH}$ .

Production of Sodium Hydroxide

The raw material for producing sodium hydroxide is  $\text{NaCl}$ . Sodium hydroxide is produced by the electrolysis of concentrated aqueous solution of sodium chloride called brine.

When electricity is passed through a concentrated solution of  $\text{NaCl}$  called brine it decomposes

## Science Formulas for class 8 by Entrancei

5. Reaction with non metals:  
Reaction with oxygen gives  $\text{CO}_2$  gas and reaction with  $\text{Cl}_2$  gas gives phosgene.

$$\begin{array}{lcl} 2\text{CO} + \text{O}_2 & \longrightarrow & 2\text{CO}_2 \\ \text{CO} + 2\text{H}_2 & \longrightarrow & \text{CH}_3\text{OH} \text{ Methyl Alcohol} \\ \text{CO} + \text{Cl}_2 & \longrightarrow & \text{COCl} \text{ Carbonyl Chloride (phosgene)} \\ \text{CO} + \text{S} & \longrightarrow & \text{COS} \text{ Carbonyl Sulphide} \end{array}$$

6. As reducing agent/oxidizing agent:

CO reduces oxides of less active metals.

$$\begin{array}{lcl} \text{Fe}_3\text{O}_4 + 3\text{CO} & \longrightarrow & 2\text{Fe} + 3\text{CO}_2 \\ \text{CuO} + 3\text{CO} & \longrightarrow & \text{Cu} + \text{CO}_2 \\ \text{ZnO} + 3\text{CO} & \longrightarrow & \text{Zn} + \text{CO}_2 \end{array}$$

7. Biochemical process:  
Combine with haemoglobin of blood to form carboxyhaemoglobin. This makes CO gas poisonous.

## USE OF CARBON MONOXIDE

1. Carbon monoxide can be used as a fuel as it produces a large amount of heat on combustion. The fuels like coal gas, water gas and producer gas contain carbon monoxide as an important constituent.  
2. Carbon monoxide is used in the extraction of metals like nickel from the mixture of other metals.  
3. In the manufacture of chemicals.  
4. It behaves as a reducing agent in the extraction of metals.  
5. In the preparation of war gas like phosgene.

## CARBON DIOXIDE

## PHYSICAL PROPERTIES OF CARBON DIOXIDE

Colour	: Colourless
Odour	: No characteristic smell
Taste	: Slightly sour in taste
Vapour density	: Heavier than air $\text{V.D.} = 22$
Solubility	: Highly soluble in water
Nature	: Acidic. Changes blue litmus to red.
Boiling point	: $-52^\circ\text{C}$
Special feature	: Not Poisonous. Its solid form is called dry ice which sublimates. It can also be easily liquefied at 70 atm pressure at room temperature.

## CHEMICAL PROPERTIES OF CARBON DIOXIDE

## 209 GENERAL CHARACTERISTIC PROP. OF BASES :

1) Bitter Taste

2) Slippery or Soapy touch.

3) Turns RED Litmus BLUE.

4) Turns PHENOLPHTHALEIN solution PINK.

5) CORROSIVE NATURE : Some bases like caustic soda ( $\text{NaOH}$ ), caustic potash ( $\text{KOH}$ ) etc are corrosive.

6) CONDUCTION OF ELECTRICITY : Like acids, the sol. of bases in water also conduct electricity.

7) REACTION WITH ACIDS (Neutralisation)

$\boxed{\text{BASE} + \text{ACID} \rightarrow \text{SALT} + \text{WATER}}$

8) REACTIONS WITH METALS : Some bases react with bases to liberate hydrogen gas alongwith the formation of a salt. Ex:

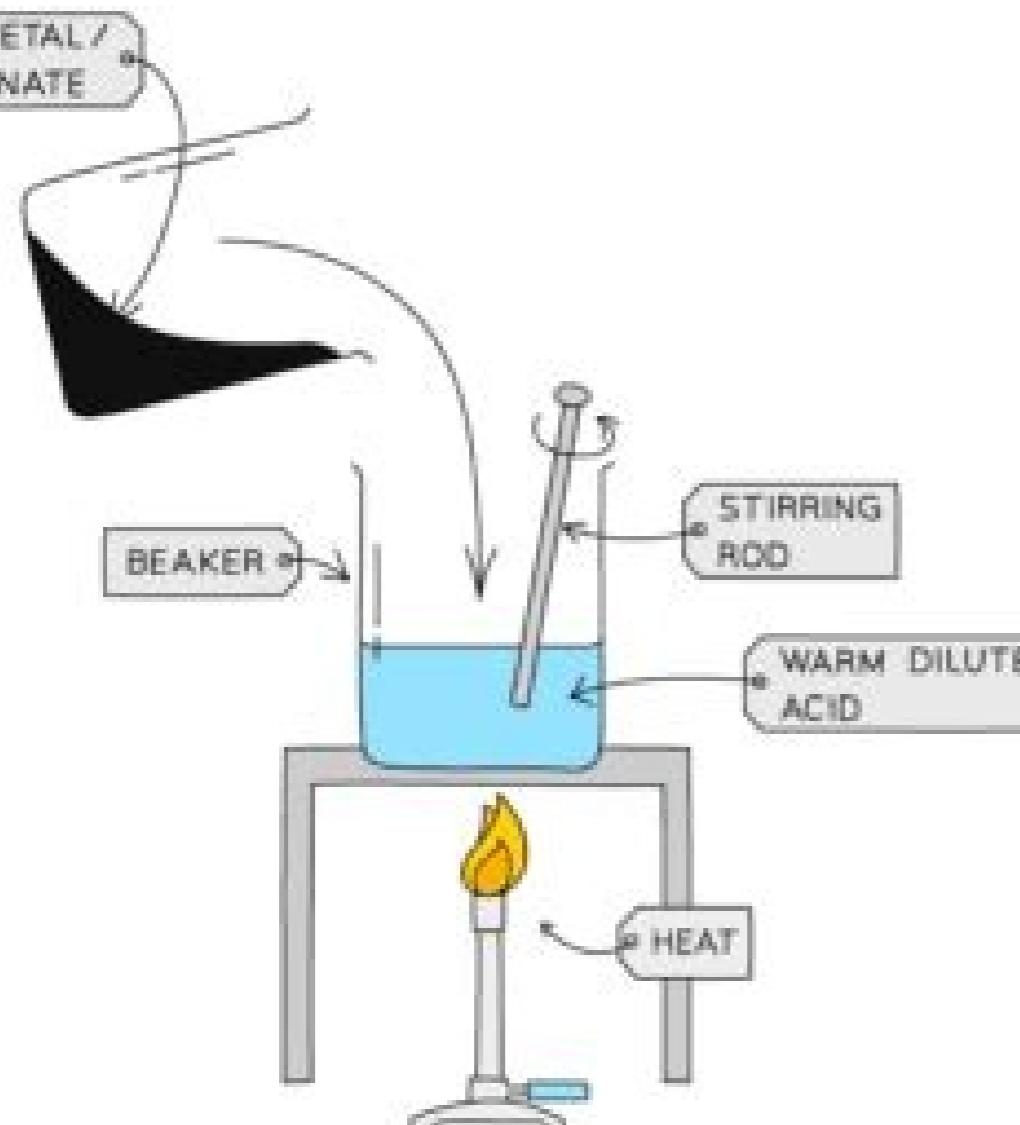
$2\text{NaOH} + \text{Zn} \rightarrow \text{Na}_2\text{ZnO}_2 + \text{H}_2$   
(Sodium Zincate)

$2\text{NaOH} + 2\text{Al} \rightarrow 2\text{NaAlO}_2 + 3\text{H}_2$   
(Sodium Aluminate)

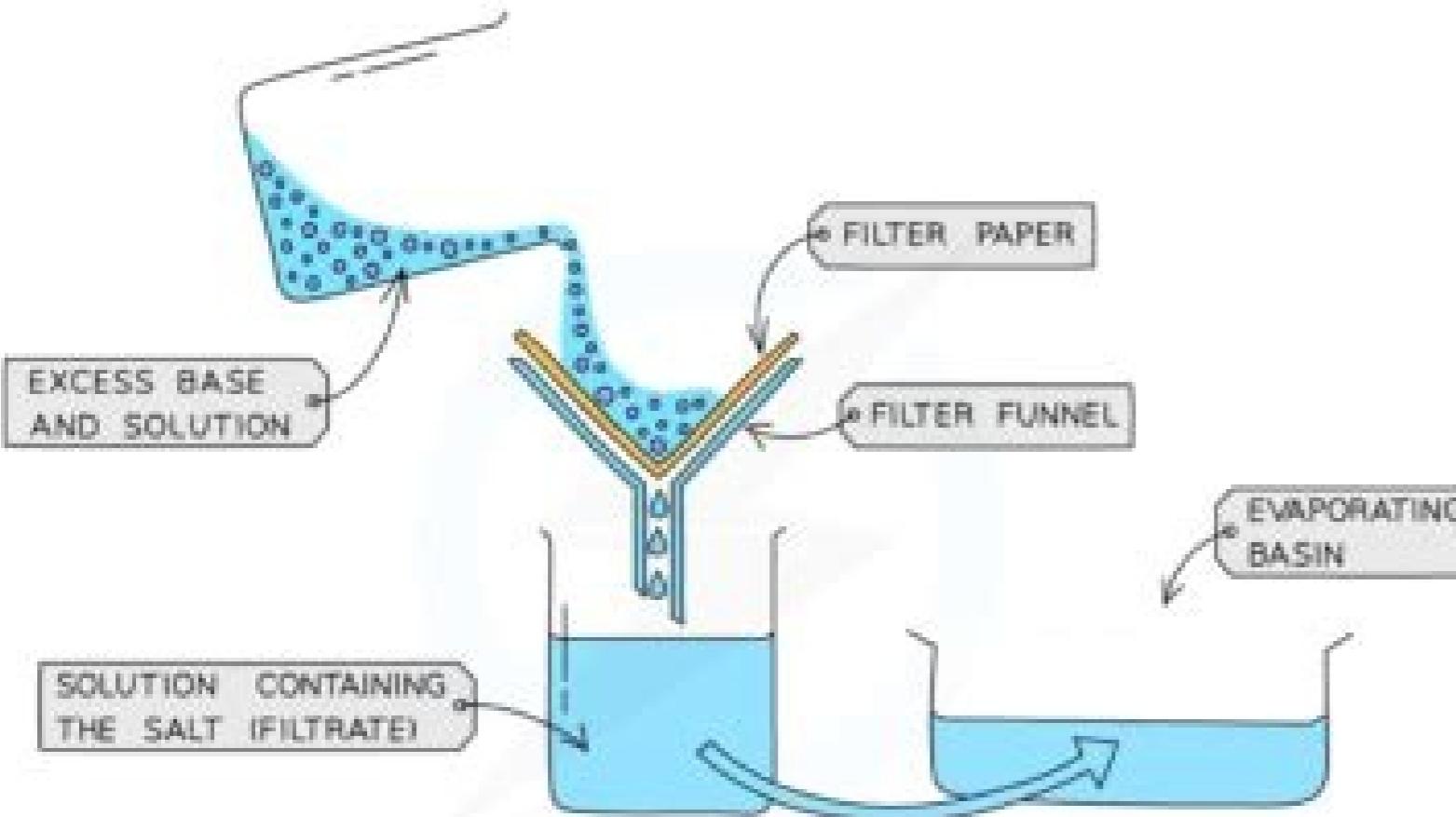
9) REACTION WITH NON METAL OXIDES

$\boxed{\text{Base} + \text{Non metal Oxide} \rightarrow \text{Salt} + \text{Water}}$

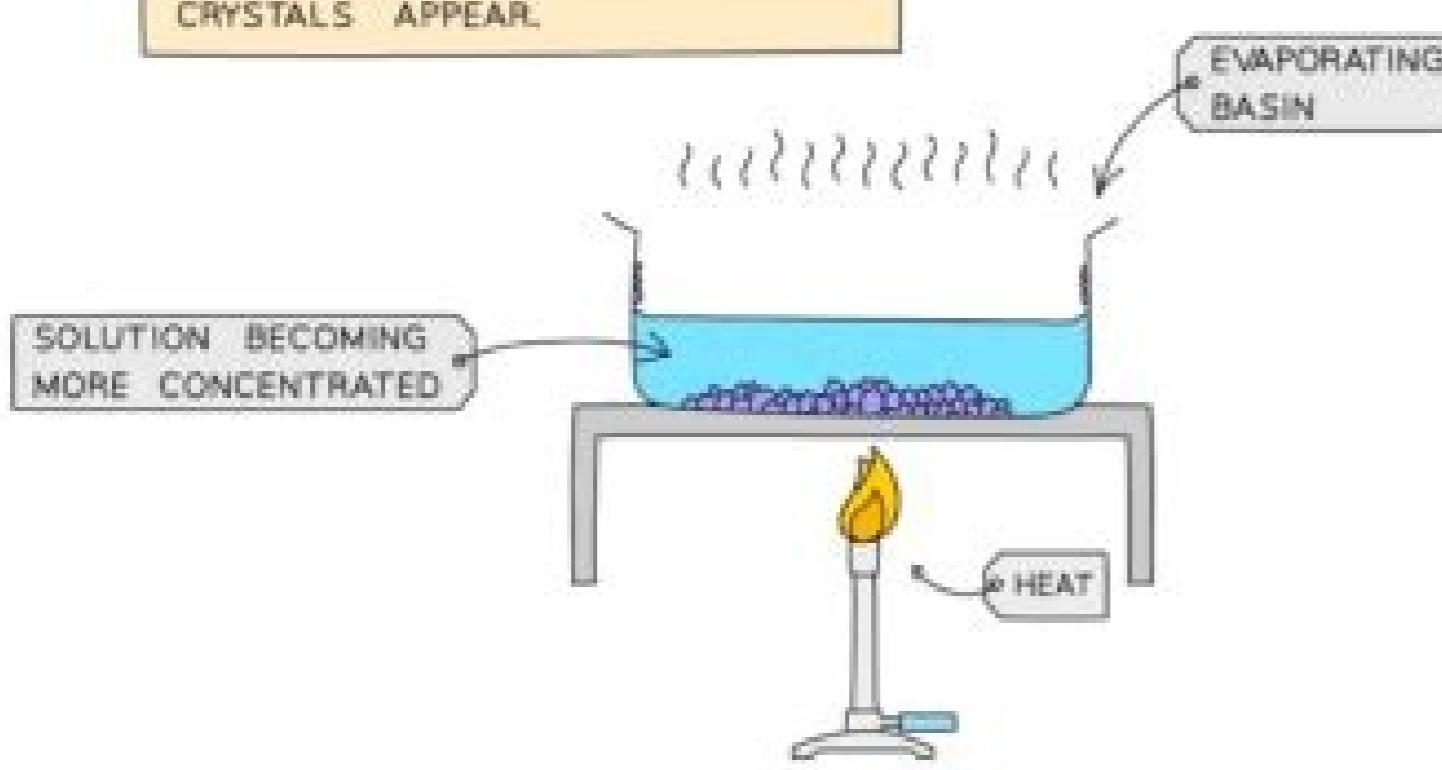
HEAT ACID UNTIL WARM, THEN ADD  
METAL / BASE / CARBONATE, STIRRING CONSTANTLY  
UNTIL IT STOPS DISAPPEARING



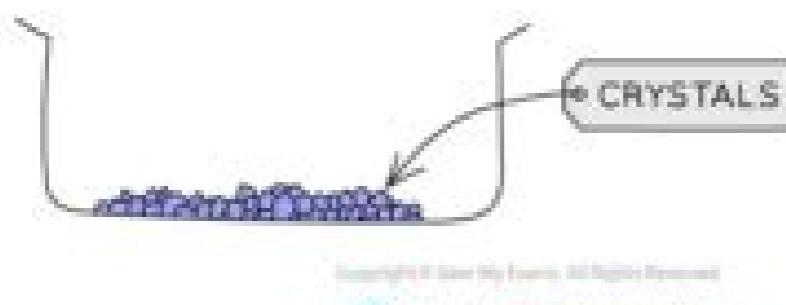
FILTER MIXTURE TO REMOVE EXCESS BASE.  
TRANSFER SOLUTION TO EVAPORATING BASIN.



EVAPORATE WATER FROM SOLUTION  
USING A BUNSEN BURNER UNTIL  
CRYSTALS APPEAR.



REMOVE EVAPORATING BASIN FROM HEAT  
AND ALLOW FILTRATE TO DRY AND CRYSTALLIZE



Types of mineral salts found in the human body include sodium, chloride, fluorine, iodine, zinc, and iron. There is a pH range on which the indicator is useful. At high pH, the concentration of  $\text{H}_3\text{O}^+$  is low, so the indicator displays the balance towards the right side of the equation and the color B. Over the course of history, the salts derived from these areas have been used in areas of medicine, food, healing, cleaning and more. Mineral salts of the human body are very popular for human medicinal uses. If the indicator is a weak base, the base and its conjugate acid show different colors. An example of a weak acid indicator is phenolphthalein, which is colorless as a weak acid but dissociates into water to form a magenta or red-violet anion. They are also used to evaluate pH values and for interesting demonstrations of color science. The indicators are chosen, then mixing some drops with a solution will produce a color which may be associated with a nominal value of the approximate pH. Soaking in salts helps to relax muscles, soothe pains and pains, lower inflammation, helps to eliminate harmful toxins from the body of your body, increase circulation and promote quality healthy for body and mind. Another aspect of the health of the mineral salts which are also beneficial for human hair. For a weak acid indicator with the formula of General Hin, equilibrium is reached in the solution according to the chemical equation:  $\text{HIN}(\text{AQ}) + \text{H}_2\text{O}(\text{L}) \rightleftharpoons \text{IN}^-(\text{AQ}) + \text{H}_3\text{O}^+(\text{AQ})$ . The acid, which is a color other than the base in- (AQ). In the absence of these critical mineral salts, humans may be more willing<sup>1</sup> to develop health problems such as anemia, gouters, and general fatigue. Food uses mineral salts are abundant in many foods and beverages, which constitute an important part of human diet. So, which might be good for a solution could be a scarce choice to test another solution. Some indicators cannot actually identify acids or bases, but I can only tell you the pH of an acid or a base. For example, the orange methyl works only at an acid pH. More from reference.com The species and its married have different colors. The point where an indicator changes colors is different for each chemical substance. Health and wellness One of the most popular uses of mineral salts is in the aromatherapy area for health and wellness. The salt is also found in waterfalls, springs, seas, oceans and other water bodies with a high mineral content. A few PA and water. Geology Geology Assistaion, U.S. department Indoor. An acid-base indicator is a weak acid or a weak base that shows a color change when the concentration of hydrogen ions ( $\text{H}^+$ ) or hydroxide ions ( $\text{OH}^-$ ) changes into an aqueous solution. Different plants and household chemicals can be used as pH indicators, but in a laboratory context, these are the most common chemicals used as indicators: color acid color base color PH range Thymol Blue (First Change) Red Yellow 1 , 2 - 2.8 1.5 orange methyl Red Yellow 3.2 - 4.4 3.7 Bromocroton green yellow blue 3.8 - 5.4 4.7 Methyl red yellow 4.8 - 6.0 5.1 Bromothymyl blue yellow blue 6.0 - 7.6 7.0 Red yellow phenol red 6.8- 8.4 7.9 Thymol Blue ( Second Change) Yellow Blue 8.0 - 9.6 8.9 Phenolphthalein Colorless Magenta 8.2 -10.0 9.4 The "ACID" and "base" colors are relative. The mineral salts clean a person's scalp and are a natural defense to prevent issues caused by dryness. A basic solution has a pH more than 7, while an acid solution has a pH less than 7. Mineral salts are important to support and maintain body digestion, fluid balance, muscle function and blood pressure, between the others. The red cabbage can also be used as an acid-base indicator. For example, a popular sea salt is the salt that comes from Dead. Another the pink salt of the Hovalayan, which comes from land and sea. Why is it that minerals found in salts are an essential part of the human body? human? helping him work the right way. Examples of mineral salts include sodium, chlorine, calcium, phosphate, magnesium, sulfur, potassium and phosphorus. Minerals are important for humanity for their use in food and in ways that the human body needs to develop and function normally. Aqueous solutions with a pH of 7 are considered neutral. Acid base indicators are substances used to approximately determine where a solution falls on the pH scale. The equilibrium constant for the reaction can be determined using the equation:  $K_{\text{IN}} = \frac{[\text{H}_3\text{O}^+][\text{IN}^-]}{[\text{HIN}]}$  where  $K_{\text{IN}}$  is the dissociation constant of the indicator. This crystallized substance is refined in table salt or used as rock salt in other natural forms and is used in health and well-being, food preparation and human balance. History since the early days of humanity, natural salts have been extracted by humans from the bodies of water and the earth. One of the most common minerals in the world is salt. Mineral salts are also found in carbonated drinks, beer, cheeses, food seasonings, bread, canned foods and cooked products. Because the acidity and alkalinity these refer to pH, they can also be known as pH indicators. The pH indicators of the acid base indicators include Litmus paper, phenolphthalein and red cabbage juice. Acid-base indicator is a weak acid or a weak base that dissociates in water to produce the weak acid and its conjugate base or the weak base and its conjugate acid. You will find mineral salts in a variety fruit, vegetables, eggs, red meat, poultry, fish and grains. In an acidic solution, the balance is left, then the solution is colorless (too small magenta anion to be visible), but as the pH increases, the balance moves to the right and the magenta color is visible. The acid-base indicators rep rep elaznesse "alovat ad elas li, oidos id orurolc ad etnemlapicnirp otitutsoC .anilacla o elartuen ,adica "asouqca enoizulos anu es eranimreted rep itazzilitu icimihi ittodor and human life. The color change occurs at the point where the concentration of the acid and anion base are equal:  $[\text{HIn}] = [\text{In}^-]$  which is the point where half of the indicator is in acid form and the other half is its conjugate base. When the pH is low, the concentration of the hydronium ion  $\text{H}_3\text{O}^+$  is high and equilibrium is toward the left, producing the color A. Also, note some popular indicators display more than one color change as the weak acid or weak base dissociates more than once. If the indicator is a weak acid, the acid and its conjugate base are different colors. In chemistry and cooking, many substances dissolve in water to make it either acidic or basic/alkaline. It would be the same color above a certain pH (acidic) and also at neutral and alkaline values. Pexels Mineral salts are substances extracted from bodies of water and below the Earth's surface. A particular type of acid-base indicator is a universal indicator, which is a mixture of multiple indicators that gradually changes color over a wide pH range. Also Known As: pH indicator Perhaps the best-known pH indicator is litmus. Thymol Blue, Phenol Red, and Methyl Orange are all common acid-base indicators. It's made by combining sodium and chlorine to form sodium chloride. Generally, all underground salts come from salt mines, salt flats, salt fields, and similar

mineral-rich areas. Acid-base indicators are most often used in a titration to identify the endpoint of an acid-base reaction. Mineral salts are also used in food preservation and food for animals. The most commonly used mineral salt is table salt, which is used all around the world. Using mineral salts can help build up one's hair volume, shine, and overall strength. Combined with essential oils in a tub filled with water, mineral salts provide a balm of relaxation to the user.

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