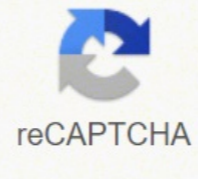




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Open



Uses of Common Salt :-

- + It is used as a raw material for making useful chemicals. e.g. NaOH (caustic soda) Na_2CO_3 (washing soda), $NaHCO_3$ (Baking Soda),
- + In cooking food. It improves the flavour of food.
- + It is required by our body for the working of nervous system.
- + It is used as a preservative in pickles meat + fish.
- + It is used in the manufacture of soap.
- + It is used to melt ice which collects on the road in cold countries.

Chemicals From Common Salt

Sodium Hydroxide :-

Sodium hydroxide is commonly known as caustic soda. The chemical formula of sodium hydroxide is NaOH.

Production of Sodium Hydroxide

The raw material for producing sodium hydroxide is NaCl. Sodium hydroxide is produced by the electrolysis of concentrated aqueous solution of sodium chloride called brine

When electricity is passed through a concentrated solution of NaCl called brine it decomposes

Science Formulas for class 8 by Entrance!

5. Reaction with non metals:
Reaction with oxygen gives CO₂ gas and reaction with Cl₂ gas gives phosgene.

Reaction with oxygen gives CO₂ gas and reaction with Cl₂ gas gives phosgene.

2CO + O₂ → 2CO₂
 CO + 2H₂ → CH₃OH Methyl Alcohol
 CO + Cl₂ → COCl₂ Carbonyl Chloride (phosgene)
 CO + S → COS Carbonyl Sulphide

6. As reducing agent/oxidizing agent:
CO reduces oxides of less active metals.

Fe₂O₃ + 3CO → 2Fe + 3CO₂
 CuO + 3CO → Cu + 3CO₂
 ZnO + 3CO → Zn + 3CO₂

7. Biochemical process:
Combine with haemoglobin of blood to form carboxyhaemoglobin. This makes CO gas poisonous.

USE OF CARBON MONOXIDE

- Carbon monoxide can be used as a fuel as it produces a large amount of heat on combustion. The fuels like coal gas, water gas and producer gas contain carbon monoxide as an important constituent.
- Carbon monoxide is used in the extraction of metals like nickel from the mixture of other metals.
- In the manufacture of chemicals.
- It behaves as a reducing agent in the extraction of metals.
- In the preparation of war gas like phosgene.

CARBON DIOXIDE

PHYSICAL PROPERTIES OF CARBON DIOXIDE

Colour	: Colourless
Odour	: No characteristic smell
Taste	: Slightly sour in taste
Vapour density	: Heavier than air V.D. = 22
Solubility	: Highly soluble in water
Nature	: Acidic. Changes blue litmus to red.
Boiling point	: -22°C
Special feature	: Not Poisonous. Its solid form is called dry ice which sublimates.

It can also be easily liquefied at 70 atm pressure at room temperature.

CHEMICAL PROPERTIES OF CARBON DIOXIDE

299 GENERAL CHARACTERISTIC PROP. OF BASES :

- Bitter Taste
- Slippery or soapy touch.
- Turns RED Litmus BLUE.
- Turns PHENOLPHTHALEIN solution PINK.
- Corrosive Nature :** Some bases like caustic soda (NaOH), caustic potash (KOH) etc are corrosive.
- Conduction of electricity :** Like acids, the sol. of bases in water also conduct electricity.
- REACTION WITH ACIDS (Neutralisation)**
 $BASE + ACID \rightarrow SALT + WATER$
- REACTIONS WITH METALS :** Some bases react with bases to liberate hydrogen gas alongwith the formation of a SALT. Ex:
 $2NaOH + Zn \rightarrow Na_2ZnO_2 + H_2$ (Sodium Zincate)
 $2NaOH + 2Al \rightarrow 2NaAlO_2 + 3H_2 + 2H_2O$ (Sodium Aluminate)
- REACTION WITH NON METAL OXIDES**
 $Base + Non\ metal\ oxide \rightarrow Salt + Water$

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